

Type #21

Write a balanced equation and indicate the reaction type.

1. Aluminum nitrate combines with sodium hydroxide to produce aluminum hydroxide and sodium nitrate.



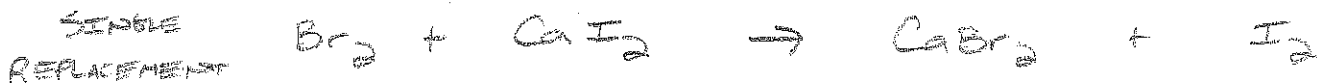
2. Potassium chlorate breaks down to form potassium chloride and oxygen.



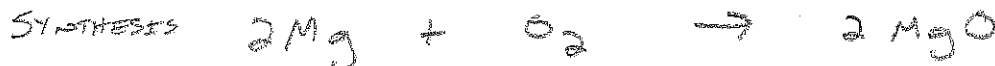
3. Ammonia reacts with oxygen to produce mononitrogen monoxide and water.



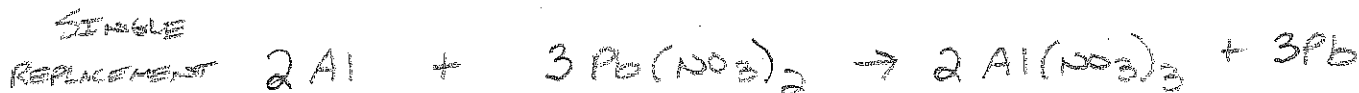
4. Calcium bromide and iodine are produced when bromine reacts with calcium iodide.



5. Magnesium reacts with oxygen to produce magnesium oxide.



6. Use a balanced equation to show the products formed when aluminum reacts with lead(II) nitrate.



7. Iron(III) oxide and water are produced when iron(III) hydroxide decomposes.



8. Diphosphorus pentoxide reacts with water to produce phosphoric acid.



9. Silver oxide reacts to produce silver and oxygen.



10. Cadmium phosphate reacts with ammonium sulfide yielding cadmium sulfide ammonium phosphate.

DOUBLE REPLACEMENT

