

Naming Acids and Bases

Name the following acids and bases:

ide
ate -ic
ite -ous

- 1) NaOH Sodium Hydroxide
- 2) H_2SO_3 Sulfurous Acid
- 3) H_2S Hydro-sulfuric Acid
- 4) H_3PO_4 Phosphoric Acid
- 5) NH_3 Ammonia
- 6) HCN Cyanic Acid
- 7) Ca(OH)_2 Calcium Hydroxide
- 8) $\text{Fe}^{+3}(\text{OH})_3^{-3}$ Iron (III) Hydroxide
- 9) H_3P Hydrophosphoric Acid

Write the formulas of the following acids and bases:

- 10) hydrofluoric acid HF
- 11) hydroselenic acid H_2SeO_4
- 12) carbonic acid H_2CO_3
- 13) lithium hydroxide LiOH
- 14) nitrous acid HNO_2
- 15) cobalt (II) hydroxide $\text{Co}^{+2}(\text{OH})_2$
- 16) sulfuric acid H_2SO_4
- 17) beryllium hydroxide $\text{Be}(\text{OH})_2$
- 18) hydrobromic acid HBr

Review: Naming Ionic Compounds

Write the formulas of the following ionic compounds (1 pt each):

- 1) ⁺² iron (II) ⁻³ arsenide Fe_3As_2
- 2) ⁺² lead (II) sulfate $PbSO_4$
- 3) ⁺⁴ lead (IV) hydroxide $Pb(OH)_4$
- 4) ⁺² copper (II) acetate $Cu(C_2H_3O_2)_2$
- 5) ⁺² beryllium ⁻¹ chloride $BeCl_2$
- 6) ammonium chromate $(NH_4)_2CrO_4$
- 7) ⁺¹ silver oxide Ag_2O
- 8) potassium sulfide K_2S

Write the names of the following ionic compounds (1 pt each):

- 9) KI Potassium Iodide
- 10) ⁺² ⁻² $Mn_2(SO_3)_7$ Manganese (VII) Sulfite
- 11) ⁺² ⁻⁴ $SnBr_4$ Tin (IV) Bromide
- 12) ⁺² Mg_3P_2 Magnesium Phosphide
- 13) NaF Sodium Fluoride
- 14) ⁺² $Sr(MnO_4)_2$ Strontium Permanganate
- 15) ⁺³ ⁻³ $Cr(PO_4)_2$ Chromium (III) Phosphate
- 16) ⁺³ Al_2Se_3 Aluminum Selenide