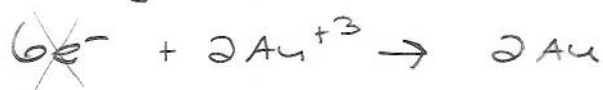
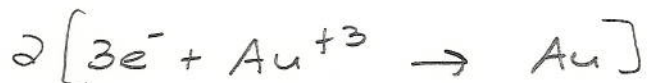
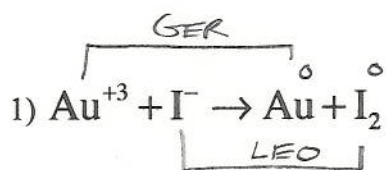


## Review for Redox Quiz



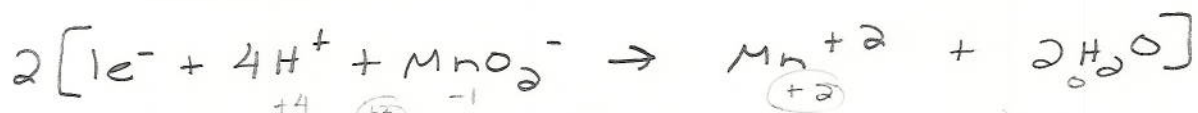
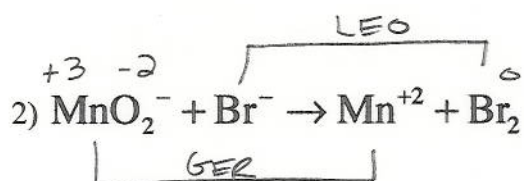
+6

-6

0

0

$$0 = 0 \checkmark$$



-2

+8

-2

0

+4

0

$$+4 = +4 \checkmark$$

1 - What happens during reduction?

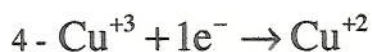
An atom gains electrons.

2 - What happens during oxidation?

An atom loses electrons.

3 - What does "redox" mean?

Shortcut for oxidation/reduction - transfer of electrons.



a. Does the above half-reaction show oxidation or reduction?

b. How do you know?

Oxidation # of Cu went down because it gained electrons.

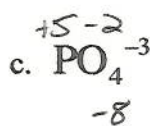
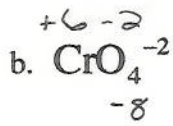
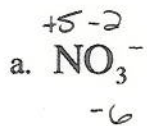


a. Does the above half-reaction show oxidation or reduction?

b. How do you know?

Oxidation # of Al went up because it lost electrons.

6 - Determine the oxidation number of each element the following:



7 - Balance the charge of the following reaction.

