

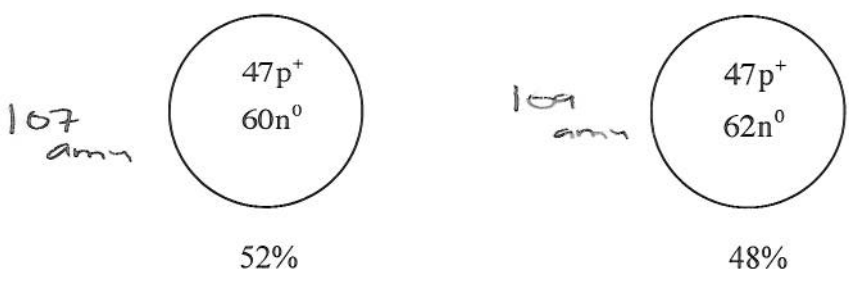
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Sub-Atomic Review

1) Complete the table below.

Atomic #	Mass #	Number of Protons	Number of Neutrons	Number of Electrons	Symbol of Element
53	127	53	74	56	I ⁻³
35	80	35	45	35	Br

2) Silver has two isotopes, as seen below with their percent abundance. Determine the average atomic mass of Silver.



$$\text{Ag } 107 \text{ amu} \times .52 = 55.64$$

$$\text{Ag } 109 \text{ amu} \times .48 = \overset{+}{52.32}$$

$$\text{Ag } 107.96 \text{ amu}$$