

Chemistry Worksheet  
Balancing Equations/Reaction Types → Following

Name \_\_\_\_\_  
Date \_\_\_\_\_ Period \_\_\_\_\_

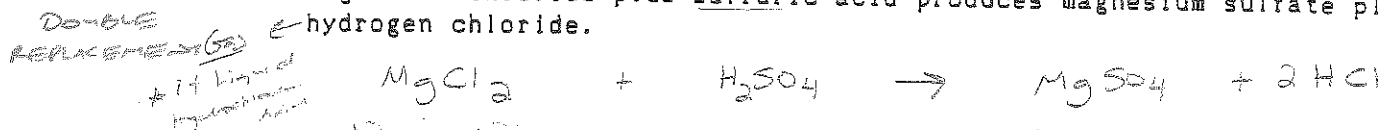
#19 / #20  
↓  
Type  
of  
Rxn

Rewrite each of the following word equations as balanced chemical equations. For each equation, state the type of reaction (single replacement, double replacement synthesis, decomposition) that best fits.

1. Sulfur plus oxygen yields sulfur dioxide.



2. Magnesium chloride plus sulfuric acid produces magnesium sulfate plus hydrogen chloride.



3. Calcium carbonate decomposes to form calcium oxide plus carbon dioxide.



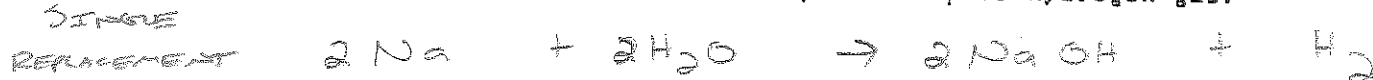
4. Nitrogen reacts with hydrogen to produce ammonia (NH<sub>3</sub>).



5. Hydrochloric acid combines with sodium hydroxide to form sodium chloride and water.



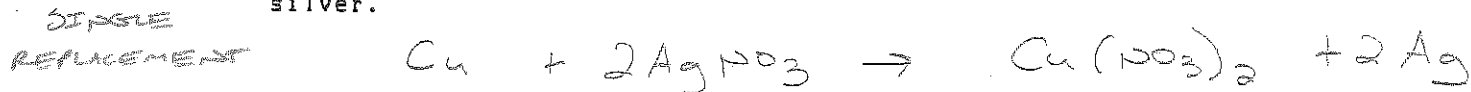
6. Sodium plus water yields sodium hydroxide plus hydrogen gas.



7. Lead (II) oxide produces lead and oxygen.



8. Copper combines with silver nitrate to produce copper (II) nitrate and silver.



9. Phosphoric acid decomposes into diphosphorus pentoxide and water.



10. Calcium chloride reacts with silver nitrate to form silver chloride and calcium nitrate.



11. Sodium oxide added to water yields sodium hydroxide.



12. Sodium bromide plus chlorine combine to form sodium chloride and bromine

