

Chemistry Worksheet
Balancing Equations

Easier - Formulas
already here!!

Name _____
Date _____ Period _____

- #10
1. $\underline{2} \text{NaClO}_3 \rightarrow \underline{2} \text{NaCl} + \underline{3} \text{O}_2$
2. $\underline{1} \text{Co(OH)}_3 + \underline{3} \text{HNO}_3 \rightarrow \underline{1} \text{Co(NO}_3)_3 + \underline{3} \text{HOH}$
3. $\underline{2} \text{ZnS} + \underline{3} \text{O}_2 \rightarrow \underline{2} \text{ZnO} + \underline{2} \text{SO}_2$
4. $\underline{2} \text{Al} + \underline{6} \text{HCl} \rightarrow \underline{2} \text{AlCl}_3 + \underline{3} \text{H}_2$
5. $\underline{1} \text{Al}_2(\text{SO}_4)_3 + \underline{1} \text{Ca}_3(\text{PO}_4)_2 \rightarrow \underline{2} \text{AlPO}_4 + \underline{3} \text{CaSO}_4$
6. $\underline{2} (\text{NH}_4)_3\text{PO}_4 + \underline{3} \text{Ba(OH)}_2 \rightarrow \underline{1} \text{Ba}_3(\text{PO}_4)_2 + \underline{6} \text{NH}_4\text{OH}$
7. $\underline{3} \text{Fe} + \underline{4} \text{H}_2\text{O} \rightarrow \underline{1} \text{Fe}_3\text{O}_4 + \underline{4} \text{H}_2$
8. $\underline{1} \text{H}_2\text{O} + \underline{1} \text{N}_2\text{O}_3 \rightarrow \underline{2} \text{HNO}_2$
9. $\underline{2} \text{PbO}_2 \rightarrow \underline{2} \text{PbO} + \underline{1} \text{O}_2$
10. $\underline{1} \text{Cl}_2 + \underline{2} \text{LiI} \rightarrow \underline{2} \text{LiCl} + \underline{1} \text{I}_2$
11. $\underline{2} \text{Al} + \underline{3} \text{Pb(NO}_3)_2 \rightarrow \underline{2} \text{Al(NO}_3)_3 + \underline{3} \text{Pb}$
12. $\underline{1} \text{Na}_2\text{SO}_3 + \underline{2} \text{HCl} \rightarrow \underline{2} \text{NaCl} + \underline{1} \text{H}_2\text{O} + \underline{1} \text{SO}_2$
13. $\underline{2} \text{K} + \underline{2} \text{H}_2\text{O} \rightarrow \underline{2} \text{KOH} + \underline{1} \text{H}_2$
14. $\underline{1} \text{BaCO}_3 + \underline{1} \text{C} + \underline{1} \text{H}_2\text{O} \rightarrow \underline{2} \text{CO} + \underline{1} \text{Ba(OH)}_2$
15. $\underline{3} \text{HF} + \underline{2} \text{N}_2 \rightarrow \underline{1} (\text{HF})_3\text{N}_4$
16. $\underline{1} \text{MnO}_2 + \underline{4} \text{HCl} \rightarrow \underline{1} \text{MnCl}_2 + \underline{1} \text{Cl}_2 + \underline{2} \text{H}_2\text{O}$
17. $\underline{1} \text{Cu} + \underline{2} \text{HNO}_3 \rightarrow \underline{1} \text{Cu(NO}_3)_2 + \underline{1} \text{HNO}_2 + \underline{1} \text{H}_2\text{O}$
18. $\underline{1} \text{Te} + \underline{1} \text{H}_2\text{O} \rightarrow \underline{1} \text{TeO} + \underline{1} \text{H}_2$
19. $\underline{2} \text{Ac(OH)}_3 \rightarrow \underline{1} \text{Ac}_2\text{O}_3 + \underline{3} \text{H}_2\text{O}$
20. $\underline{1} \text{Ca}(\text{AlO}_2)_2 + \underline{8} \text{HCl} \rightarrow \underline{2} \text{AlCl}_3 + \underline{1} \text{CaCl}_2 + \underline{4} \text{H}_2\text{O}$